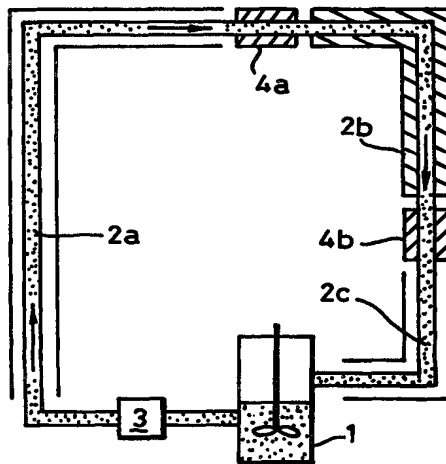




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(54) Title: METHOD AND COMPOUND FOR INHIBITING THE PLUGGING OF CONDUITS BY GAS HYDRATES

**(57) Abstract**

A method is disclosed for inhibiting the plugging of a conduit for the transport of hydrocarbon fluids by gas hydrates, wherein use is made of a hydrate formation inhibitor component of formula (I): $(R_1)(R_2)(R_3)(R_4)A^+Y^-$, wherein two of R_1-R_4 are independently normal or branched alkyls having 4 or 5 carbon atoms, two of R_1-R_4 are independently representing organic moieties having at least 8 carbon atoms, A represents a nitrogen or a phosphorus atom, Y represents an anion and wherein at least one of R_1-R_4 represents a $-(CH_2-CHR_5-O)_p-(CH_2)_q-(CHR_6-CH_2)_r-(CH_2-CHR_7)_s-(CHR_8)_t-O-C(O)-R^9$ moiety with the meaning as defined in the description. Compositions containing a hydrate formation inhibitor component according to formula (I) and a hydrocarbon having from one to eight hydrocarbons as well as the hydrate formation inhibitor components, such as the di-esters of di-butyl -di-isopropanol (or -di-isobutanol) ammonium bromide and coconut fatty acid are believed to be novel.

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METHOD AND COMPOUND FOR INHIBITING THE PLUGGING OF
CONDUITS BY GAS HYDRATES

Background of the Invention

This invention relates to a method and compound for inhibiting the plugging by gas hydrates of conduits containing a mixture of low-boiling hydrocarbons and water. In particular, it relates to a method and compound for continuing the flow of a mixture of hydrocarbons and water through a conduit after a shutdown which caused separation of hydrate crystals from the mixture.

The problem of the formation of gas hydrates (clathrates of gases in a lattice consisting of water molecules) is well known in the art. Low-boiling hydrocarbons, in particular methane, may under conditions of elevated pressure and reduced temperature tend to form gas hydrate crystals with water present in natural gas or crude oil.

Such gas hydrate crystals, when allowed to form and grow inside a conduit such as a pipeline, tend to block or even damage the conduit. A number of methods has been suggested to prevent such blocking of which the use of crystal growth inhibitors is considered to be very attractive.

In International Patent Application Publication WO 96/34177 the use of a class of hydrate formation inhibitors is described having the general formula $(R_1)(R_2)(R_3)(R_4)A^+Y^-$ wherein two of R_1-R_4 are independently normal or branched alkyls having 4 or 5 carbon atoms, two of R_1-R_4 are independently representing organic moieties having at least eight carbon atoms, A represents a nitrogen or a phosphorus atom and Y represents an anion.

Attractive results have been obtained when using the di-ester of di-butyl-di-ethanol ammonium bromide and coconut fatty acid, both with respect to decreasing the temperature at which hydrates tend to be formed (as described in experiment A 3b of said patent specification) and with respect to restarting the flow of the medium after stopping (as described in experiment B 1b of said specification). In the experiment B 1b it is described that after a temperature decrease of 11 °C, followed by stopping the circulation of the medium used which resulted in the slow separation of a layer of very fine hydrate crystals from the hydrate forming medium, the circulation could be restarted whereby the layer of loose powder hydrates became readily resuspended into the hydrocarbon liquids resulting in the formation of the hydrate suspension which was observed prior to shutdown.

One of the problems still to be solved in managing the flow of suspended gas hydrates in conduits is being able to restart the flow after an interruption, either scheduled or unscheduled, when powder hydrates have been formed under rather severe conditions. Even with the preferred di-ester of di-butyl-di-ethanol ammonium bromide and coconut fatty acid restarts at rather severe conditions appear to be problematic.

Summary of the Invention

It has now surprisingly been found that a specific group of esters, in addition to having the attractive properties as described in WO 96/34177 also exhibits an improved restart performance when subjected to rather severe conditions which makes them very attractive for such duties.

The present invention thus relates to a method for inhibiting the plugging of a conduit, the conduit containing a mixture comprising an amount of hydrocarbons

having from one to eight carbon atoms and an amount of water wherein the amounts of hydrocarbons and water could form hydrates at conduit temperatures and pressures, the method comprising the steps of:

5 adding to the mixture an amount of a hydrate formation inhibitor component of formula



wherein two of R_1 - R_4 are independently

normal or branched alkyls having 4 or 5 carbon atoms,

10 two of R_1 - R_4 are independently representing organic moieties having at least 8 carbon atoms,

A represents a nitrogen or phosphorus atom,

Y represents an anion

and wherein at least one of R_1 - R_4 represents a

15 $-(CH_2-CHR_5-O)_p-(CH_2)_q-(CHR_6-CH_2)_r-(CH_2-CHR_7)_s-$
 $(CHR_8)_t-O-C(O)-R_9$ moiety wherein R_5 - R_8 each independently represent a hydrogen atom or a C_1 - C_4 alkyl group and at least one of R_6 - R_8 is not a hydrogen atom; p represents 0 or an integer of up to 50; q, r, s and t are 0, 1 or 2
20 and together at least 1 and not more than 4 and wherein when q represents 1 or 2 the sum of r, s and t is at least 1 and, wherein when t and the sum of q, r, s and t represent 2, R_8 does not represent a methyl group; and R_9 represents an (cyclo)alkyl, alkenyl, aryl, arylalkyl,
25 arylalkenyl, alkylaryl or alkenylaryl group having at least a carbon chain of 6 atoms,

the amount of hydrate formation inhibitor component being effective to inhibit formation of hydrates in the mixture at conduit temperatures and pressures, and

30 flowing the mixture containing the hydrate formation inhibitor component through the conduit.

Preferably, A represents nitrogen and two of R_1 - R_4 independently contain between 8 and 20 carbon atoms, preferably in the range from 10 to 16 carbon atoms.

35 Suitably, compounds are used wherein two of R_1 - R_4 contain

the same number of carbon atoms, each being at least 8. Preferably, use can be made of components wherein two of R_1 - R_4 represent cocoyl moieties (i.e. the alkyl chain present in coconut fatty acids or similar compounds).

5 Preference is given to compounds wherein R_6 , R_7 and R_8 independently represent hydrogen or a methyl or ethyl group.

Preferred ammonium compounds in accordance with the present invention are those wherein p is zero, the sum of
10 q , r and s is 2 and R_9 represents an alkyl or alkenyl group of between 9 and 18 carbon atoms. Very good results have been obtained with compounds for which q and t are 1 and R_8 represents a methyl or ethyl group.

Suitably, the anion (Y^-) represents a hydroxide, a
15 carboxylate, a halide such as chloride or bromide, a sulphate or an organic sulphonate. Preferably, Y^- represents a chloride, bromide or sulphate.

The presence of the carboxylic moiety in the compounds in accordance with the invention is of great
20 advantage because of their biodegradable properties which renders them eminently suitable for the envisaged use. A further advantage in the envisaged use is that such compounds are sparingly soluble in water which allows discarding production water containing only marginal
25 concentrations of such compounds.

The amount of hydrate formation inhibitor compound to be used lies suitably between 0.05 and 11 %wt, based on the water content of the hydrocarbon/water mixture. Preferably, the amount is between 0.1 and 5 %wt, in
30 particular between 0.1 and 0.5 %wt, based on the water content as defined.

If desired, corrosion inhibitors may be added to the hydrocarbon/water mixture. Corrosion inhibitors known to those skilled in the art can be suitably applied. If
35 desired, polymers of ethylenically unsaturated N-hetero-

cyclic carbonyl compounds can be added to the hydro-carbon/water mixture. Reference is made to the polymers or copolymers of N-vinyl-2-pyrrolidone as described in International Patent Application Publication WO 93/25798.

5 Compounds according to formula I, in which two of R_1 - R_4 are independently representing organic moieties having at least 8 carbon atoms, A represents a nitrogen or a phosphorus atom, Y represents an anion, and wherein at least one of R_1 - R_4 represents a

10 $-(CH_2-CHR_5-O)_p-(CH_2)_q-(CHR_6-CH_2)_r-(CH_2-CHR_7)_s-(CHR_8)_t-O-C(O)-R_9$ moiety wherein R_5 - R_8 each independently represent a hydrogen atom or a C_1 - C_4 alkyl group and at least one of R_6 - R_8 is not a hydrogen atom; p represents 0 or an integer of up to 50; q, r, s and t are 0, 1 or 2
15 and together at least 1 and not more than 4 and wherein when q represents 1 or 2 the sum of r, s and t is at least 1 and, wherein when t and the sum of q, r, s and t represent 2, R_8 does not represent a methyl group; and R_9 represents an (cyclo)alkyl, alkenyl, aryl, arylalkyl,
20 arylalkenyl, alkylaryl or alkenylaryl group having at least a carbon chain of 6 atoms, are believed to be novel.

Suitable compounds are those in which A represents nitrogen and two of R_1 - R_4 independently contain between 8
25 and 20 carbon atoms, preferably in the range from 10 to 16 carbon atoms. Further suitable compounds are those in which two of R_1 - R_4 contain the same number of carbon atoms, each being at least 8.

Preference is given to compounds in which two of
30 R_1 - R_4 represents cocoyl moieties (i.e. the alkyl chain present in coconut fatty acids or similar compounds). In particular, preference is given to compounds in which R_6 , R_7 and R_8 independently represent hydrogen, a methyl group or an ethyl group.

Preferred ammonium compounds are those in which p is zero, the sum of q, r, s and t is 2 and R₉ represents an alkyl or alkenyl group of between 9 and 18 carbon atoms; in particular compounds in which q and t are 1 and R₈ represents a methyl or ethyl group whilst r and s are zero.

Suitably, the anion (Y⁻) represents a hydroxide, a carboxylate, a halide (such as chloride or bromide), a sulphate or an organic sulphonate. Preferably, Y⁻ represents a chloride, bromide or sulphate.

Most preferred compounds are the di-ester of di-butyl di-isopropanol ammonium bromide and coconut fatty acid and the di-ester of di-butyl di-isobutanol ammonium bromide and coconut fatty acid.

Example

The following Example will illustrate the invention.
Description of equipment

In the experiments that are detailed below field conditions were simulated by using a high-pressure flow loop facility which is schematically shown in Figure 1 and which consists of a stainless steel pipelooop (2a-c) having an inner diameter of 19 mm and an effective length of 108 metres, a mixing tank (1) and a gear pump (3) for circulating a hydrate forming mixture of water and liquid hydrocarbons through the loop. The pipelooop can be seen as being divided into 9 sections (each having a length of 12 metres) and each of which is equipped with a thermometer and a differential pressure meter allowing the monitoring of the pressure drop over each individual section.

Sections 1-6 (2a) and section 9 (2c) are surrounded by a coaxial pipe through which a temperature-controlled liquid is circulated in counter-flow to the hydrate forming medium (which flows from section 1 to section 9).

Sections 7 and 8 (2b) are thermally well insulated and equipped with viewing windows (mounted near the inlet of section 7 (4a) and the outlet of section 8 (4b) to allow the visual observation of the hydrate forming medium in the pipeloop.

Hydrate formation is triggered by cooling 1 cm² of the inner surface of the pipeloop near the end of section 3 to a constant temperature of -15 °C. This "cold spot" was switched-off immediately after the first hydrates were formed.

Standard filling and pre-conditioning procedure

In all experiments described hereafter, the loop facility (having a total volume of 62.5 litres) was (at a temperature of 24 °C) initially filled with 5 litres of water, 39.2 litres of a hydrocarbon liquid, such as SHELLSOL D60 (trade name for a mixture of paraffinic and naphthenic hydrocarbons, mainly in the C₁₀-C₁₂ range, available from Shell Oil Company, Houston, Texas) and 3.2 kilograms of propane. Subsequently, methane was introduced until the equilibrium pressure was 78 bara. This procedure leads to the formation of a three-phase system (i.e. a vapour phase, a liquid aqueous phase and a liquid hydrocarbon phase) in which can form stable hydrates at temperatures below 19 °C. In all experiments the liquid phases of the hydrate forming medium were circulated through the pipe loop at a rate of 120 grams/second (or 540 litres per hour) which corresponds to a Reynolds number of approximately 8000 (turbulent flow). Prior to the start of each experiment the hydrate forming medium was circulated for approximately one day at a temperature of 23 °C to obtain thermodynamic equilibrium and an even distribution of the liquid phases throughout the entire system.

The effect of an additive (hydrate forming inhibitor component) was assessed by comparing the experimental results of a blank test (in which no additive had been added to the hydrate forming medium) with those of an additive test (in which the system was doped with the additive concerned) and which test was carried out under the same conditions of the blank test.

The experiments represent the hydrate recirculation mode in which the temperature of the hydrate forming medium is kept constant throughout the entire test facility.

Preparation of active compounds:

I. Preparation of the di-cocoyl ester of di-butyl di-isopropanol ammonium bromide:

a) Preparation of di-butyl amino-propanol-2

In a vessel 516 g dibutylamine (ex Aldrich) and 296 g butanol (ex Merck) were mixed and heated under nitrogen to 100 °C. The pressure was 1 bar. About 100 g propylene oxide was added causing the pressure to rise to 2 bar. A further amount of 132 g propylene oxide was added causing the pressure to rise to about 3 bar.

Thereafter the total mixture was heated to 120 °C. This caused the pressure to rise to 5 bar which dropped to 3.3 bar after 90 minutes. The reaction mixture was kept overnight and then subjected to flash distillation at reducing pressure. This yielded 498 g of (dibutyl-amino) propanol-2 (purity > 95%).

b) Preparation of dibutyl di-2-hydroxypropyl ammonium bromide

234 g of the product obtained under a) was mixed with 468 g water, 234 g isopropanol and 169 g hydrogen bromide (48% wt) and heated under nitrogen to 70 °C. 58 g propylene oxide was added in the course of a few minutes, causing an increase in the pressure from 1.64 bar to

2.03 bar. After 2 hours the pressure had become constant at 1.85 bar.

The crude product was worked up by flashing off the solvent followed by 2x desalting in isopropanol. Non-polar organics were removed by dissolving the sample in water, raising the pH to 14 and extracting the water layer twice with diethylether. After readjustment of the pH with hydrogen bromide the solution was subjected to evaporation to dryness. A polar organic impurity, propylene glycol, was removed by slurrying with diethylether. The final amount of propylene glycol in the otherwise pure product (yield 145 g) amounted to less than 3 %mole.

c) Preparation of the dicocoylester of dibutyl di-isopropanol ammonium bromide

Product (48.8 g) obtained under b) and 164 g cocoyl anhydride were stirred in Rotavap equipment at 100 °C. The ammonium salt is solid, does not melt and does not dissolve under the prevailing conditions. After stirring for 2 hours at 100 °C a nearly homogeneous mixture had been formed. According to $^1\text{H-NMR}$ a complete conversion had taken place.

By means of wiped film evaporation (120 °C/ 4.10^{-2} mbar) 147 g of crude quaternary ammonium compound were obtained. Since the crude compound still contained some higher molecular weight carboxylic acids and anhydrides (preventing crystallisation of the ammonium compound from common solvents) it was subjected to a treatment with acetic anhydride (stirring for 1 hour at ambient temperature). After filtration of some precipitate, the liquid product was concentrated at 50 °C/ 3.10^{-2} mbar using Rotavap equipment. After subjecting it to a further purification step by means of wiped film evaporation (100 °C at 3.10^{-2} mbar) the

quaternary ammonium compound was obtained in 75% purity as characterised by ^1H -NMR and ^{13}C -NMR, the remainder consisting of anhydrides. Precipitation from 3 litres of pentane yielded 50 g of the pure dicocoyl ester of di-2-hydroxypropyl ammonium bromide.

^{13}C -NMR data: 13 ppm:2C; 14 ppm:2C; 19 ppm:2C; 20 ppm:2C; 22 ppm:2C; 24 ppm:2+2C; 29 ppm:ca 14C; 32 ppm:2C; 34 ppm:2C; 62 ppm:2C; 63 ppm:2C; 65 ppm:2C and 172 ppm:2C.

^1H -NMR: 0.9 ppm: CH_3 (ex coco), 6H; 1.0 ppm: CH_3 (ex butyl), 6H; 1.2-2.1 ppm:div. CH_2 , ca 46H; 2.3 ppm: CH_2CO , 4H; 3.2-3.4 ppm: $\text{C}_3\text{H}_7\text{CH}_2\text{N}$, 4H; 3.6-4.0 and 4.2-4.4 ppm: $\text{NCH}_2\text{CH}(\text{CH}_3)\text{O}$, 4H; 5.3-5.5 ppm: CHOCO , 2H.

When a sample of the compound produced was subjected to saponification under standard conditions it was found that the acids present in the product were the same as initially present in the starting cocoyl compounds.

II. Preparation of the dicocoyl ester of dibutyl di-isobutanol ammonium bromide.

a) Preparation of di-butyl amino-butanol-2

In a 1.7 litre Medinex autoclave were mixed 310 g of di-butyl amine (ex Aldrich), 134 g butanol and 125 g butylene oxide (ex Aldrich). The autoclave was flushed 4 times with 5 bar nitrogen. The mixture is heated to 120 °C and kept under stirring for 16 hours at this temperature. After cooling down the contents of the autoclave are subjected to distillation using a Vigreux column. The product fraction (analysed using Gas Chromatography) amounted to 306 g (yield: 88%).

b) Preparation of di-butyl-di-2-butanol ammonium bromide

In a 1.7 litre Medinex autoclave were mixed 251 g of product obtained under a), 400 g of water, 300 g of isopropanol, 83 g butylene oxide (ex Aldrich) and 169 g hydrogen bromide (48% wt). The autoclave was flushed

4 times with 5 bar nitrogen. The mixture is heated to 80 °C and kept under stirring for 48 hours at this temperature. After cooling down the contents of the autoclave, 0.5 litres of isopropanol/water is removed using Rotavap equipment. Sodium hydroxide is added to obtain pH 14. After addition of 500 ml water, the product is extracted twice with 300 ml diethylether to remove amines present. The aqueous layer is then acidified using hydrogen bromide until pH 1-2. Thereafter water is removed by distillation (80 °C/vacuum). Still remaining water is removed using azeotropic distillation with toluene.

The residue is stirred in isopropanol and filtrated to remove inorganic salts. The residue is then subjected to evaporation to dryness (yielding about 150 g). By stirring the product with 500 ml diethylether, the viscous liquid mixture slowly transforms into a crystalline solid in diethylether. After filtration, the solid is stirred with 500 ml diethylether, filtrated again and dried under high vacuum. The diethylether fraction contained 24 g butanediol.

The solid is then stirred in 300 ml of CH₂Cl₂ and filtrated again to remove the last remnants of inorganic salts. By concentrating the filtrate a viscous oil is obtained. After a final treatment with 500 ml diethylether, the solid reappeared and was subjected to filtration and drying. The yield was 116 g (29%) and the structure confirmed with both ¹H-NMR and ¹³C-NMR.

c) Preparation of the dicocoylester of di-butyl di-isobutanol ammonium bromide

Coco anhydride (150 g) and di-butyl di-2-butanol ammonium bromide obtained in step b) (50 g) were mixed in Rotavap equipment and heated to 100 °C. At that temperature the product from step b) has melted causing a two

layer system. After continuous stirring for one hour a homogeneous mixture had been formed. A sample was subjected to ^1H -NMR analysis and showed nearly quantitative conversion.

5 The product obtained was subjected three times to wiped film evaporation ($120\text{ }^\circ\text{C}/10^{-1}\text{ mbar}$, then at $120\text{ }^\circ\text{C}/5.10^{-2}\text{ mbar}$ and finally at $120\text{ }^\circ\text{C}/3.10^{-2}\text{ mbar}$). The residue weighs 134 g and contains some acid, anhydride and ammonium compound.

10 Finally, the product was stirred for 1 hour with acetic acid anhydride at ambient temperature. A solid product is formed which is separated by filtration and is identified as long chain anhydride. The acetic acid anhydride and acetic acid are removed by Rotavap distillation
15 ($50\text{ }^\circ\text{C}$ and 2 mbar). The product is subjected to wiped film evaporation at $120\text{ }^\circ\text{C}$ and 4.10^{-2} mbar . The yield of the isobutanol compound amounts to 119 g (70% purity).

Experiment 1a (blank test)

20 The test facility was filled with 5 litres of water, 3.2 kilograms of propane and 39.2 litres of a mixture consisting of 85% SHELLSOL D 60 and 15% of SHELLSOL R (SHELLSOL is a trademark), after which methane was added until the equilibrium pressure at $24\text{ }^\circ\text{C}$ was 78 bar. This procedure leads to the formation of a three-phase system
25 (i.e. a vapour phase, a liquid aqueous phase and a liquid hydrocarbon phase) in which can form stable hydrates at temperatures below $19\text{ }^\circ\text{C}$. The liquid phases of the hydrate forming medium were circulated through the pileloop at a rate of 120 g/second (or 540 litres/hour)
30 which corresponds to a Reynolds number of approximately 8000 (turbulent flow). Prior to the start of each experiment the hydrate forming medium was circulated for approximately one day at a temperature of $23\text{ }^\circ\text{C}$ to obtain

thermodynamic equilibrium and an even distribution of the liquid phases throughout the entire system.

The experiment was started by cooling the hydrate forming medium at rate of 1 °C/hr. Because no heating was applied in the ninth section, the temperature of the hydrate forming medium was independent of the position of the medium in the test facility. In this type of test the hydrates which are carried by the flow become severely crushed when they pass through the gear pump. During this experiment the first increase in the pressure drop was observed after four hours by which time the temperature of the medium was 18.8 °C. The circulation could be maintained for another hour during which the pressure drop increased continuously until the loop became completely blocked by hydrates. At the time of blocking the temperature of the hydrate forming medium was 18.0 °C.

Experiment 1b

This experiment was identical to experiment 1a except for the addition of 12.5 grams of the diester of dibutyl-diethanol ammonium bromide and coconut fatty acid, giving a concentration of 0.25 %wt of hydrate formation inhibitor component. Eleven hours after the start of the cooling cycle, at which the temperature of the medium was 12 °C, the circulating liquids became hazy whereas the pressure in the system dropped rapidly indicating that a substantial amount of hydrates were formed. The cooling cycle was continued for another 11 hours after which the temperature of the hydrate forming mixture was reduced to 1 °C and only a slight increase of the pressure drop over the loop was observed. The medium was circulated for another two hours during which the pressure drop did not increase. At this stage the pressure of the system had dropped to 52 bar indicating that practically all water

was converted into hydrates. Subsequently the circulation was stopped resulting in the slow separation of a layer of very fine hydrate crystals from the hydrate forming medium. This shut-down condition was maintained for the next 22 hours during which the temperature of the medium was kept at 1 °C. When the circulation was restarted the layer of loose powder hydrates became readily resuspended into the hydrocarbon liquids resulting in the formation of the hydrate suspension which was observed prior to shut-down. Also the pressure drop over the pipeloop had not increased with respect to the situation before the circulation was stopped.

Experiment 1c

The experiment as described in experiment 1b was repeated except for allowing a drop in temperature of 14 °C. Restart of the circulation, which was smoothly possible under the conditions as described in experiment 1b, after 2.5 hours of shut down was not possible. When performing a similar test with a drop in temperature of 18 °C, a restart was hardly possible even after 1 hour after shut down.

Experiment 1d (according to the invention)

The experiment as described in experiment 1 b was repeated except for using 0.50 %wt of the diester of dibutyl diisopropanol ammonium bromide and coconut fatty acid as the hydrate formation inhibitor component. When allowing a temperature drop of 14 °C the circulation could be restarted without any problem, even after a shutdown of 16.8 hours. When the experiment was repeated allowing a temperature drop of 18 °C a smooth restart could be effected even after 20.3 hours.

Experiment 1e

The experiment as described in experiment 1b was repeated except for using 0.50 %wt of the diester of

dibutyl dipropanol ammonium bromide and coconut fatty acid. When allowing a temperature drop of 14 °C a restart, even after 0.5 hours after shut-down failed completely.

5 Experiment 1f (according to the invention)

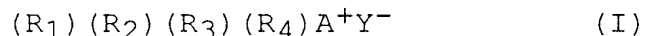
 The experiment as described in experiment 1b was repeated except for using 0.5 %wt of the diester of dibutyl diisobutanol ammonium bromide and coconut fatty acid. When allowing a temperature drop of 18 °C a smooth
10 restart could be achieved even after a shut-down of no less than 17 hours.

 From the experimental data it will be clear that the branched components used in the process according to the invention have a marked and unexpected performance
15 advantage with respect to restart of circulation after a shut-down under severe conditions.

C L A I M S

1. A method for inhibiting the plugging of a conduit,
the conduit containing a mixture comprising an amount of
hydrocarbons having from one to eight carbon atoms and an
amount of water wherein the amounts of hydrocarbons and
5 water could form hydrates at conduit temperatures and
pressures, the method comprising the steps of :

adding to the mixture an amount of a hydrate
formation inhibitor component of formula



10 wherein two of R_1 - R_4 are independently

normal or branched alkyls having 4 or 5 carbon atoms,

two of R_1 - R_4 are independently representing organic
moieties having at least 8 carbon atoms,

A represents a nitrogen or phosphorus atom,

15 Y represents an anion

and wherein at least one of R_1 - R_4 represents a

$-(CH_2-CHR_5-O)_p-(CH_2)_q-(CHR_6-CH_2)_r-(CH_2-CHR_7)_s-(CHR_8)_t-O-C(O)-R_9$ moiety wherein R_5 - R_8 each
independently represent a hydrogen atom or a C_1 - C_4 alkyl
20 group and at least one of R_6 - R_8 is not a hydrogen atom; p
represents 0 or an integer of up to 50; q, r, s, and t
are 0, 1 or 2 and together at least 1 and not more than 4
and wherein when q represents 1 or 2 the sum of r, s, and
t is at least 1 and, wherein when t and the sum of q, r,
25 s and t represent 2, R_8 does not represent a methyl
group; and R_9 represents an (cyclo)alkyl, alkenyl, aryl,
arylalkyl, arylalkenyl, alkylaryl or alkenylaryl group
having at least a carbon chain of 6 atoms,

the amount of hydrate formation inhibitor component
30 being effective to inhibit formation of hydrates in the
mixture at conduit temperatures and pressures, and

flowing the mixture containing the hydrate formation inhibitor component through the conduit.

2. The method according to claim 1, wherein A represents nitrogen and two of R_1 - R_4 independently contain between 8 and 20 carbon atoms, preferably in the range from 10 to 16 carbon atoms.

3. The method according to claim 1 or 2, wherein two of R_1 - R_4 contain the same number of carbon atoms, each being at least 8.

4. The method according to one or more of claims 1-3, wherein R_6 , R_7 and R_8 each independently represent hydrogen or a methyl or ethyl group.

5. The method according to claim 4, wherein p is zero, the sum of q, r, s, and t is 2 and R_9 represents an alkyl or alkenyl group of between 9 and 18 carbon atoms.

6. The method according to claim 5, wherein q and t are 1 and R_8 represents a methyl or ethyl group.

7. The method according to one or more of claims 1-6, wherein R_9 represents the carbon chain of coconut fatty acid or tallow fatty acid.

8. The method according to one or more of claims 1-11, wherein the anion represents a hydroxide, a carboxylate, a halide, a sulphate or an organic sulphonate.

9. The method according to one or more of claims 1-8, wherein between 0.05 and 11 %wt, based on the water content, of the hydrate formation inhibitor component is added to the water/hydrocarbon mixture.

10. The method according to claim 9, wherein the amount of hydrate formation inhibitor component added is between 0.1 and 5 %wt, preferably between 0.1 and 0.5 %wt, based on the water content.

11. The method according to one or more of claims 1-10 wherein a corrosion inhibitor is added to the hydrocarbon/water mixture.

12. The method according to one or more of claims 1-11, wherein a polymer of an ethylenically unsaturated N-heterocyclic carbonyl compound is added to the hydrocarbon/water mixture.

5 13. The method according to one or more of claims 1-12, wherein the step of flowing the mixture containing the hydrate formation inhibitor component through the conduit is effected after a shutdown which caused separation of hydrate crystals from the mixture.

10 14. The method according to claim 13, wherein the step of flowing the mixture is effected with a hydrocarbon/water mixture which has been subjected during shutdown to a decrease in temperature of more than 12 °C, in particular to a decrease of at least 14 °C.

15 15. Compounds according to the general formula I in which at least one of R₁-R₄ represents a

-(CH₂-CHR₅-O)_p-(CH₂)_q-(CHR₆-CH₂)_r-(CH₂-CHR₇)_s-
(CHR₈)_t-O-C(O)-R₉ moiety wherein R₅-R₈ each independently represent a hydrogen atom or a C₁-C₄ alkyl group and at
20 least one of R₆-R₈ is not a hydrogen atom; p represents 0 or an integer of up to 50; q, r, s and t are 0, 1 or 2 and together at least 1 and not more than 4 and wherein when q represents 1 or 2 the sum of r, s and t is at least 1, and, wherein when t and the sum of q, r, s and t
25 represent 2, R₈ does not represent a methyl group; and R₉ represents an (cyclo)alkyl, alkenyl, aryl, aralkyl, aryl-alkenyl, alkylaryl or alkenylaryl group having at least a carbon chain of 6 atoms.

30 16. Compounds according to claim 15, in which two of R₁-R₄ have the meaning as defined in claim 15.

17. Compounds according to claim 15 or 16 in which A represents nitrogen and two of R₁-R₄ independently contain between 8 and 20 carbon atoms, preferably in the range from 10 to 16 carbon atoms.

18. Compounds according to one or more of claims 15-17, in which two of R₁-R₄ contain the same number of carbon atoms, each being at least 8.

19. Compounds according to one or more of claims 15-18, in which R₆, R₇ and R₈ each independently represent hydrogen or a methyl or ethyl group.

20. Compounds according to claim 19, in which p is zero, the sum of q, r, s and t is 2 and R₉ represents an alkyl or alkenyl group of between 9 and 18 carbon atoms.

21. Compounds according to claim 20, in which q and t are 1 and R₈ represents a methyl or ethyl group.

22. Compounds according to one or more of claims 15-21, in which R₉ represents the carbon chain of coconut fatty acid or tallow oil fatty acid.

23. Compounds according to one or more of claims 15-22, in which the anion represents a hydroxide, a carboxylate, a halide, a sulphate or an organic sulphonate.

24. Di-ester of di-butyl di-isopropanol ammonium bromide and coconut fatty acid.

25. Di-ester of di-butyl-di-isobutanol ammonium bromide and coconut fatty acid.

26. A composition comprising one or more hydrate formation inhibitor components according to the general formula I, a hydrocarbon liquid and, optionally, a corrosion inhibitor, in which the components according to formula I comprise compounds in which two of R₁-R₄ are independently

normal or branched alkyls having 4 or 5 carbon atoms, two of R₁-R₄ are independently representing organic

moieties having at least 8 carbon atoms,

A represents a nitrogen or phosphorus atom,

Y represents an anion

and wherein at least one of R₁-R₄ represents a

$-(\text{CH}_2-\text{CHR}_5-\text{O})_p-(\text{CH}_2)_q-(\text{CHR}_6-\text{CH}_2)_r-(\text{CH}_2-\text{CHR}_7)_s-$

$(\text{CHR}_8)_t-\text{O}-\text{C}(\text{O})-\text{R}_9$

moiety wherein R₅-R₈ each independently represent a hydrogen atom or a C₁-C₄ alkyl group and at least one of R₆-R₈ is not a hydrogen atom; p represents 0 or an integer of up to 50; q, r, s and t are 0, 1 or 2 and together at least 1 and no more than 4 and wherein when q represents 1 or 2 the sum of r, s and t is at least 1 and, wherein when t and the sum of q, r, s and t represent 2, R₈ does not represent a methyl group; and R₉ represents an (cyclo)alkyl, alkenyl, aryl, arylalkyl, arylalkenyl, alkylaryl or alkenylaryl group having at least a carbon chain of 6 atoms.

27. Compositions according to claim 26, in which two of R₁-R₄ have the meaning as defined in claim 26.

28. Compositions according to claims 26 or 27, in which A represents nitrogen and two of R₁-R₄ independently contain between 8 and 20 carbon atoms, preferably in the range from 10 to 16 carbon atoms.

29. Compositions according to one or more of claims 26-28, in which two of R₁-R₄ contain the same number of carbon atoms, each being at least 8.

30. Compositions according to one or more of claims 26-29, in which R₆, R₇ and R₈ each independently represent hydrogen or a methyl or ethyl group.

31. Compositions according to claim 30, in which p is zero, the sum of q, r, s and t is 2 and R₉ represents an alkyl or alkenyl group of between 9 and 18 carbon atoms.

32. Compositions according to claim 26 in which q and t are 1 and R₈ represents a methyl or ethyl group.

33. Compositions according to one or more of claims 26-32, in which R₉ represents the carbon chain of coconut fatty acid or tallow oil fatty acid.

34. Compositions according to one or more of claims 26-33, in which the anion represents a hydroxide, a carboxylate, a halide, a sulphate or an organic sulphonate.

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35. Compositions according to one or more of claims 26-34, in which the hydrocarbon liquid is SHELLSOL D60 and/or SHELLSOL R.

5 36. Compositions according to one or more of claims 26-35, in which the compound according to formula I is the di-ester of di-butyl di-isopropanol ammonium bromide and coconut fatty acid.

10 37. Compositions according to one or more of claims 26-36, in which the compound according to formula I is the di-ester of di-butyl di-isobutanol ammonium bromide and coconut fatty acid.

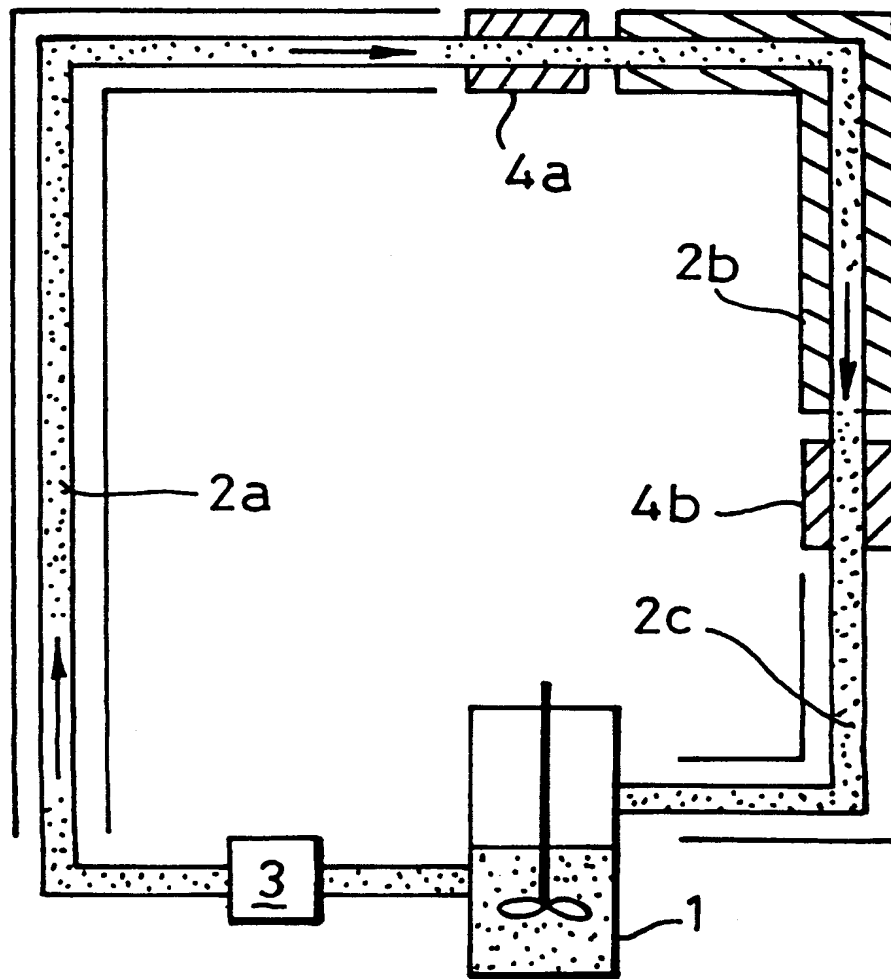


FIG. 1

INTERNATIONAL SEARCH REPORT

International Application No

PCT/EP 98/05745

A. CLASSIFICATION OF SUBJECT MATTER

IPC 6 E21B37/06 C10L3/00 F17D1/02 F15D1/00 C07C219/06

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 E21B C10L F17D F15D C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category °	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 96 34177 A (SHELL INT RESEARCH ;SHELL CANADA LTD (CA)) 31 October 1996 cited in the application see claims 1-15,18 see page 5, line 14 - line 17 see examples A,B	1-4, 7-13, 15-19, 22-30, 33-35
X	--- GOGOL, SCHAEFER: "synthese und eigenschaften photoreaktiver, amphiphiler benzophenonderivate." LIEBIGS ANN. CHEM.,1987, pages 589-596, XP002092327 see page 595 --- -/--	15-21

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

° Special categories of cited documents :

"A" document defining the general state of the art which is not considered to be of particular relevance

"E" earlier document but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

"&" document member of the same patent family

Date of the actual completion of the international search

4 February 1999

Date of mailing of the international search report

22/02/1999

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De Herdt, 0

INTERNATIONAL SEARCH REPORT

International Application No

PCT/EP 98/05745

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	KUNITAKE: "excimer formation and phase separation of hydrocarbon and fluorocarbon bilayer membranes." BULL. CHEM. SOC. JPN, vol. 56, no. 11, 1983, pages 3235-3242, XP002092328 see page 3236 -----	15-20,22
A	WO 95 17579 A (SHELL INT RESEARCH ;SHELL CANADA LTD (CA)) 29 June 1995 -----	

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

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